

इंटरनेट

मानक

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“जानने का अधिकार, जीने का अधिकार”

Mazdoor Kisan Shakti Sangathan

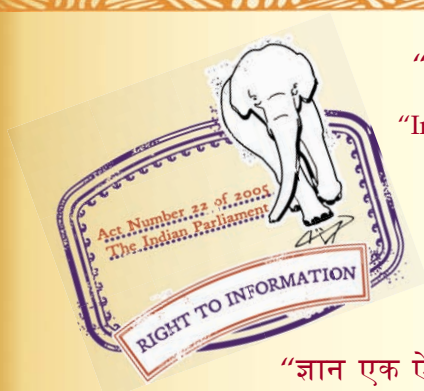
“The Right to Information, The Right to Live”

“पुराने को छोड़ नये के तरफ”

Jawaharlal Nehru

“Step Out From the Old to the New”

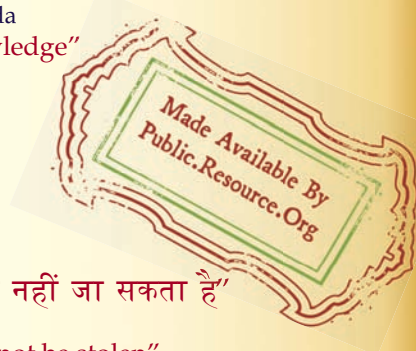
IS 3025-23 (1986): Methods of sampling and test (physical and chemical) for water and wastewater, Part 23: Alkalinity [CHD 32: Environmental Protection and Waste Management]



“ज्ञान से एक नये भारत का निर्माण”

Satyanarayan Gangaram Pitroda

“Invent a New India Using Knowledge”



“ज्ञान एक ऐसा खजाना है जो कभी चुराया नहीं जा सकता है”

Bhartrhari—Nitiśatakam

“Knowledge is such a treasure which cannot be stolen”

AMENDMENT NO. 1 SEPTEMBER 2000
TO
IS 3025 (PART 23) : 1986 METHODS OF SAMPLING
AND TEST (PHYSICAL AND CHEMICAL) FOR WATER
AND WASTEWATER
PART 23 ALKALINITY
(First Revision)

(Page 1, clause 2, line 3) — Substitute 'pH 4.5' for 'pH 3.7'.

(CHD 12)

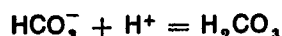
*Indian Standard***METHODS OF SAMPLING AND TEST
(PHYSICAL AND CHEMICAL) FOR
WATER AND WASTEWATER****PART 23 ALKALINITY***(First Revision)*

1. Scope — Prescribes the potentiometric and indicator methods for determination of alkalinity. These methods are applicable to determine alkalinity in water and wastewater in the range of 0.5 to 500 mg/l alkalinity as CaCO_3 . The upper range may be extended by dilution of the original sample.

2. Principle and Theory — Alkalinity of water is the capacity of that water to accept protons. It may be defined as the quantitative capacity of an aqueous medium to react with hydrogen ions to pH 8.3 (phenolphthalein alkalinity) and then to pH 3.7 (total alkalinity or methyl orange alkalinity). The equation in its simplest form is as follows:



From pH 8.3 to 3.7, the following reaction may occur:



3. Interferences — Free available residual chlorine markedly affects the indicator colour response. The addition of minimal volumes of sodium thiosulphate eliminates this interference. Substances such as salt of weak organic or inorganic acids present in large amount may interfere. Oils and greases may also interfere by coating the electrode. Coloured or turbid samples may interfere in end point. Analyse such samples by potentiometric titration.

4. Sampling and Storage — Sampling and storage shall be done as prescribed in IS:3025 (Part 1)-1986 'Methods of sampling and test (physical and chemical) for water and wastewater: Part 1 Sampling (first revision)'.

5. Sample Preparation — The sample aliquot used for analysis should be either free from turbidity or should be allowed to settle prior to analysis.

6. Apparatus**6.1 pH Meter**

6.2 Burette — 50-ml capacity.

6.3 Magnetic Stirrer Assembly**7. Reagents**

7.1 Distilled Water — Distilled water used should have pH not less than 6.0. If the water has pH less than 6.0, it shall be freshly boiled for 15 minutes and cooled to room temperature. Deionized water may be used provided that it has a conductance of less than $2 \mu\text{S}/\text{cm}$ and a pH more than 6.0.

7.2 Sulphuric Acid — Dilute 5.6 ml of concentrated sulphuric acid (relative density 1.84) to one litre with distilled water.

7.3 Standard Solution of Sulphuric Acid — 0.02 N.

7.4 Phenolphthalein Indicator — Dissolve 0.5 g of phenolphthalein in 100 ml, 1:1 (v/v), alcohol water mixture.

7.5 Mixed Indicator Solution — Dissolve 0.02 g methyl red and 0.01 g bromocresol green in 100 ml, 95 percent, ethyl or isopropyl alcohol.

Adopted 31 July 1986

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Gr 1

8. Procedure

8.1 Indicator Method — Pipette 20 ml or a suitable aliquot of sample into 100-ml beaker. If the pH of the sample is over 8.3, then add 2 to 3 drops of phenolphthalein indicator and titrate with standard sulphuric acid solution till the pink colour observed by indicator just disappears (equivalence of pH 8.3). Record the volume of standard sulphuric acid solution used. Add 2 to 3 drops of mixed indicator to the solution in which the phenolphthalein alkalinity has been determined. Titrate with the standard acid to light pink colour (equivalence of pH 3.7). Record the volume of standard acid used after phenolphthalein alkalinity.

9. Calculation — Calculate alkalinity in the sample as follows:

$$\text{Phenolphthalein alkalinity (as mg/l of CaCO}_3 \text{)} = \frac{A \times N \times 50000}{V}$$

$$\text{Total alkalinity (as mg/l CaCO}_3 \text{)} = \frac{(A+B) \times N \times 50000}{V}$$

where

A = ml of standard sulphuric acid used to titrate to pH 8.3,

B = ml of standard sulphuric acid used to titrate from pH 8.3 to pH 3.7,

N = normality of acid used, and

V = volume in ml of sample taken for test.

EXPLANATORY NOTE

Alkalinity of water or wastewater is its quantitative capacity to react with a strong acid to a designated pH. Alkalinity is significant in many uses and treatments of natural and wastewaters. Alkalinity measurements are used in the interpretation and control of water and wastewater treatment processes.

This method supersedes 13 and 14 of IS:3025-1964 'Methods of sampling and test (physical and chemical) for water used in industry.